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CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided.

1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures

2. b A covalent bond consists of (a) a shared electron.

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Chapter 1 2. Chapter 2 4. Chapter 3 6. Chapter 4 8. Chapter 5 10. Chapter 6 12. Chapter 7 14. Chapter 8 16. Chapter 9 18. Chapter 1. Question Number Answer Level 1 Head Reference for Answer Difficulty 1 C. What is an Information System?

Answers to Chapters 1,2,3,4,5,6,7,8,9 - End of Chapter ...

1 . The spectrum consists of colored lines, at least one of which (probably the brightest) is red. 3 . 3.15 m 5 . $3.233 \times 10^{-19} \text{ J}$; 2.018 eV 7 .

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6.7 Chapter Summary. To ensure that you understand the material in this chapter, you should review the meanings of the following bold terms in the following summary and ask yourself how they relate to the topics in the chapter. Chemical reactions relate quantities of reactants and products.

Chapter 6 - Quantities in Chemical Reactions - Chemistry

SECTION 6.1 Nature favors arrangements in which potential energy is minimized. For example, a boulder is less likely to balance at the top of a hill than it is to roll to the bottom of a valley. The boulder at the top of the hill is not stable. Atoms are usually not stable when isolated. They usually combine to form more stable arrangements of matter.

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CHAPTER 6 Chemical Bonding

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and what they did Rutherford proposed that all of an atom's positive charge and virtually all of its mass are in a nucleus surrounded by fast-moving electrons(his nuclear model).
Boseman Suggested that molecules could get a discrete packet of energy instead of continuous light Bohr Made the current model for atoms Said that ...

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